

10-1 Review and Reinforcement

Chemical Measurements

Define each of the following terms in complete sentences.

1. mole

2. formula mass

3. formula unit

4. Avogadro's number

5. molar mass

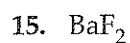
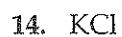
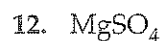
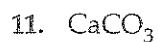
If the statement is true, write "true." If it is false, change the underlined word or words to make it true.

- _____ 6. The atomic mass of an atom is based on a comparison with the mass of sulfur-32.
- _____ 7. A mole of any element contains 6.02×10^{23} atoms.
- _____ 8. A mole of an ionic compound contains 6.02×10^{23} molecules.
- _____ 9. The molar mass of a substance in grams is always equal to the formula mass in amu.
- _____ 10. One mole of HNO_3 contains 3 atoms of oxygen.

Name _____ Date _____ Class _____

10-1 Review and Reinforcement (continued)

Find the formula mass of each of the following compounds. Show all your work.



Find the molar mass of each of the following compounds. Show all your work.

