

## 10-1 Practice Problems

---

1. What is the atomic mass of zinc?
2. Methylene chloride ( $\text{CH}_2\text{Cl}_2$ ) is used as a solvent in paint strippers. What is the formula mass of methylene chloride?
3. What is the atomic mass of nitrogen?
4. Ammonia ( $\text{NH}_3$ ) is a common household cleaning agent. What is the formula mass of ammonia?
5. Sodium hypochlorite ( $\text{NaClO}$ ) is the active ingredient in household bleach. What is the formula mass of sodium hypochlorite?
6. What is the atomic mass of neon?
7. Sodium chloride ( $\text{NaCl}$ ) is common table salt. What is the formula mass of sodium chloride?
8. What is the atomic mass of uranium?
9. Nitric acid ( $\text{HNO}_3$ ) is a strong acid. What is the formula mass of nitric acid?
10. What is the atomic mass of silver?

**10-1 Practice Problems (continued)**

Calculate the molar mass for each of the following compounds.

11. methylamine ( $\text{CH}_3\text{NH}_2$ )
12. benzene ( $\text{C}_6\text{H}_6$ )
13. propane ( $\text{C}_3\text{H}_8$ )
14. ethandiol ( $\text{C}_2\text{H}_6\text{O}_2$ )
15. chloroform ( $\text{CHCl}_3$ )
16. silver chloride ( $\text{AgCl}$ )
17. sodium hydroxide ( $\text{NaOH}$ )
18. copper(II) sulfate ( $\text{CuSO}_4$ )
19. octane ( $\text{C}_8\text{H}_{18}$ )
20. tribromosilane ( $\text{Br}_3\text{HSi}$ )